

OFFICE OF NAVAL RESEARCH

GRANT: N00014-93-1-0331

TECHNICAL REPORT No. 75

96PRO-1032

The Oxidation State of Ni in the Nickel Oxide Electrode and Related Nickel Oxide
Compounds: II. Geometric Evidence

H. Sambe, T.M. Nabi, and D.E. Ramaker

Department of Chemistry, George Washington University
Washington, DC 20052

Prepared for Publication in
Journal of Physical Chemistry

January 1997

Reproduction in whole, or in part, is permitted for any purpose of the
United States Government

This document has been approved for public release and sale; its distribution is unlimited.

19970226 034

DTIC QUALITY INSPECTED

REPORT DOCUMENTATION PAGE

Form Approved
OMB No. 0704-0188

Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing the collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden, to Washington Headquarters Services, Directorate for Information Operations and Reports, 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302, and to the Office of Management and Budget, Paperwork Reduction Project (0704-0188), Washington, DC 20503.

1. AGENCY USE ONLY (Leave blank)		2. REPORT DATE 1/97	3. REPORT TYPE AND DATES COVERED Technical	
4. TITLE AND SUBTITLE The oxidation state of Ni in the nickel oxide electrode and related nickel oxide compounds: II. Geometric evidence			5. FUNDING NUMBERS Grant #: N00014-93-1-0331	
6. AUTHOR(S) H. Sambe, T.M. Nabi, and D.E. Ramaker				
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) Department of Chemistry The George Washington University Washington, D.C. 20052			8. PERFORMING ORGANIZATION REPORT NUMBER Technical Report #75	
9. SPONSORING/MONITORING AGENCY NAME(S) AND ADDRESS(ES) Office of Naval Research 800 N. Quincy Street Arlington, VA 22217-5000			10. SPONSORING/MONITORING AGENCY REPORT NUMBER 96PRO-1032	
11. SUPPLEMENTARY NOTES				
12a. DISTRIBUTION/AVAILABILITY STATEMENT Approved for public release; distribution is unlimited.			12b. DISTRIBUTION CODE Unlimited	
13. ABSTRACT (Maximum 200 words) The electronic structures of the nickel oxide electrode (NOE), $\text{NiO}_x(\text{OH})_{2-x}$ for $x=0-1.6$ and the related nickel oxide materials including the $\text{NiO} - \text{LiNiO}_2 - \text{NiO}_2$ series and the " Ni^{4+} " compounds, $\text{KNiO}_2(\text{IO}_4)$ and BaNiO_3 , have been investigated utilizing geometric data such as the Ni-OH and Ni-O bond lengths published in the literature. The electronic structures deduced from these bond lengths agree with those found using spectroscopic data, which were presented in the preceding paper. In addition, the observed shortening of the Ni-O bond length in the $\text{NiO} - \text{LiNiO}_2 - \text{NiO}_2$ series is shown to arise from covalent shortening. The implications of this newly proposed electronic structures are as follows: (i) the charging reaction for the NOE can be viewed as $\text{OH}^- (\text{Ni}^{2+}) \rightarrow \text{O}^- (\text{Ni}^{2+}) + \text{H}^+ + \text{e}^-$, where (Ni^{2+}) reminds us that the OH^- and O^- anions are attached to the Ni^{2+} ion, which remain unchanged; (ii) the electron conduction in the charged NOE is in the oxygen layer; and (iii) the coloration of the NOE upon charging is due to the $3d (\text{Ni}^{2+}) \rightarrow 2p (\text{O}^-)$, metal-to-oxygen, charge transfer in the charged NOE.				
14. SUBJECT TERMS Nickel Hydroxides, nickel electrode, oxidation state			15. NUMBER OF PAGES 39	
			16. PRICE CODE	
17. SECURITY CLASSIFICATION OF REPORT Unclassified	18. SECURITY CLASSIFICATION OF THIS PAGE Unclassified	19. SECURITY CLASSIFICATION OF ABSTRACT Unclassified	20. LIMITATION OF ABSTRACT Unclassified	

The oxidation state of Ni in the nickel oxide electrode and related nickel oxide compounds: II. Geometric evidence

H. Sambe, T.M. Nabi, and D.E. Ramaker

Chemistry Department, George Washington University, Washington, D.C. 20052

The electronic structures of the nickel oxide electrode (NOE), $\text{NiO}_x(\text{OH})_{2-x}$ for $x=0-1.6$ and the related nickel oxide materials including the $\text{NiO} - \text{LiNiO}_2 - \text{NiO}_2$ series and the " Ni^{4+} " compounds, $\text{KNiO}_2(\text{IO}_4)$ and BaNiO_3 , have been investigated utilizing geometric data such as the Ni-OH and Ni-O bond lengths published in the literature. The electronic structures deduced from these bond lengths agree with those found using spectroscopic data, which were presented in the preceding paper. In addition, the observed shortening of the Ni-O bond length in the $\text{NiO} - \text{LiNiO}_2 - \text{NiO}_2$ series is shown to arise from covalent shortening. The implications of this newly proposed electronic structures are as follows: (i) the charging reaction for the NOE can be viewed as $\text{OH}^- (\text{Ni}^{2+}) \rightarrow \text{O}^- (\text{Ni}^{2+}) + \text{H}^+ + \text{e}^-$, where (Ni^{2+}) reminds us that the OH^- and O^- anions are attached to the Ni^{2+} ion, which remain unchanged; (ii) the electron conduction in the charged NOE is in the oxygen layer; and (iii) the coloration of the NOE upon charging is due to the $3d (\text{Ni}^{2+}) \rightarrow 2p (\text{O}^-)$, metal-to-oxygen, charge transfer in the charged NOE.

I. INTRODUCTION

In the preceding paper [referred here as paper I], we have analyzed the Ni 2p XPS (x-ray photoemission spectroscopy) and the Ni K-edge XAS (x-ray absorption spectroscopy) data for the nickel oxide electrode (NOE) and related nickel oxide compounds and found that these data consistently favor the more covalent electronic structure over the conventional ionic electronic structure for these materials, suggesting that oxidation of the Ni^{2+} ion is harder than oxidation of the O^{2-} anion in the NOE and related compounds. In this paper, we shall analyze the geometric data, namely the variations in the Ni-OH and Ni-O bond lengths.

The systems studied and notations used are the same as those in the paper I. Briefly, the systems analyzed are (a) the NOE materials, $\text{NiO}_x(\text{OH})_{2-x}$ for $x=0-1.6$; (b) the $\text{NiO} - \text{LiNiO}_2 - \text{NiO}_2$ series; and (c) the " Ni^{4+} " compounds, $\text{KNiO}_2(\text{IO}_4)$ and BaNiO_3 . A common property of all of these compounds is that the Ni atom is coordinated octahedrally by oxygen atoms. All of the compounds, except the " Ni^{4+} " compounds, are known to have a layered structure [1-3]. Here, OS.Ni will be used for oxidation state of Ni and IOS.Ni for the ionic OS.Ni, which is defined as the mean OS.Ni obtained assuming the normal oxidation states of Li^+ , K^+ , Ba^{2+} , O^{2-} , OH^- , and IO_4^- . This IOS.Ni is uniquely determined when composition of the compound is given. Therefore most of the figures in this paper and paper I are plotted as a function of this IOS.Ni.

At the end of this paper, we will summarize all evidence available including evidence presented in paper I and the others and review their consistency. After establishing the electronic structures for the NOE and the related nickel oxide compounds, we will discuss the implications of the newly proposed electronic structures.

II. Ni-O AND Ni-OH BOND LENGTHS

In this paper, the Ni-O distances of the Ni-OH and Ni-O bonds shall be referred to as the Ni-OH and Ni-O bond lengths respectively. The Ni-O distances in the NOE [i.e., $\text{NiO}_x(\text{OH})_{2-x}$] means the Ni-OH and Ni-O bond lengths in the NOE and the mean Ni-O distance for the NOE means the averaged Ni-O distance of the Ni-OH and Ni-O bond lengths in the NOE. Thus, the term "the Ni-O distance" in this paper has a restricted meaning.

Figure 1 plots the Ni-OH (open circle) and Ni-O (closed circle) bond lengths as a function of the IOS.Ni for the nickel oxide electrode materials, $\text{NiO}_x(\text{OH})_{2-x}$ [α -, β - $\text{Ni}(\text{OH})_2$ ($x=0$), β - $\text{NiO}(\text{OH})$ ($x=1$), and γ -" NiOOH " ($x=1.5$)], along with the Ni-O bond lengths for the IOS.Ni \equiv 4 compounds [i.e., $\text{Li}_{0.06}\text{NiO}_2$, $\text{KNiO}_2(\text{IO}_4)$, and BaNiO_3]. The Ni-O distances for the NOE materials are obtained by averaging the recent results reported by O'Grady [4] and Mansour [5], which agree with each other to within 0.01 Å. Those for $\text{Li}_{0.06}\text{NiO}_2$ [3], $\text{KNiO}_2(\text{IO}_4)$ [5], and BaNiO_3 [6] are obtained from the cited references. Needless to say, β - $\text{NiO}(\text{OH})$ and γ -" NiOOH " have both Ni-OH and Ni-O bonds, whereas the $\text{Ni}(\text{OH})_2$ and the IOS.Ni \equiv 4 compounds have only one kind of bond, Ni-OH or Ni-O bonds respectively. Figure 1 shows that the Ni-OH bond lengths of β - $\text{NiO}(\text{OH})$ and γ -" NiOOH " are almost equal to those of α - and β - $\text{Ni}(\text{OH})_2$ and the Ni-O bond lengths of these are close to those of the IOS.Ni \equiv 4 compounds. In short, the Ni-OH and Ni-O bond lengths are almost invariant against the IOS.Ni.

Here we analyze this invariance on alternative premises of OS.Ni=2 or OS.Ni=IOS.Ni. Let us start with the OS.Ni=IOS.Ni premise. This premise for $\text{NiO}_x(\text{OH})_{2-x}$ suggests the $\text{Ni}^{\text{IOS.Ni}+}\text{O}^{2-x}(\text{OH})^-_{2-x}$ electronic structure or the $\text{Ni}^{\text{IOS.Ni}+}\text{OH}^-$ and $\text{Ni}^{\text{IOS.Ni}+}\text{O}^{2-}$ bonds, where IOS.Ni=2+x. Therefore, the electrostatic interaction for these bonds will vary with the IOS.Ni. The $\text{Ni}^{\text{IOS.Ni}+}\text{OH}^-$ and $\text{Ni}^{\text{IOS.Ni}+}\text{O}^{2-}$ bond

lengths can be estimated rather accurately using the ionic radii for a six coordinated Ni ion proposed by Shannon [7], i.e., O^{2-} (1.26 Å), OH^- (1.23 Å), Ni^{2+} (0.83 Å), low-spin Ni^{3+} (0.70 Å), and low-spin Ni^{4+} (0.62 Å). The resultant Ni-OH (dashed line) and Ni-O (dotted line) bond lengths are inserted in Fig. 1. Let us now consider the OS.Ni=2 premise. The OS.Ni=2 premise for $NiO_x(OH)_{2-x}$ suggests the $Ni^{2+}O_x(OH)^{2-x}$ electronic structure or the $Ni^{2+}-OH^-$ and $Ni^{2+}-O^-$ bonds, independent of x and IOS.Ni ($=2+x$). Consequently, the Ni-OH and Ni-O bond lengths are expected to be invariant against the IOS.Ni. This invariance for the OS.Ni=2 assignment is included in Fig. 1 by solid lines. The experimental results agree with the solid lines (OS.Ni=2) but not with the dashed and dotted lines (OS.Ni=IOS.Ni) and hence strongly favor the OS.Ni=2 assignment for the compounds, β -NiO(OH), γ -"NiOOH", $Li_{0.06}NiO_2$, $KNiO_2(IO_4)$, and $BaNiO_3$. This assignment is consistent with the OS.Ni=2 conclusion deduced from the XPS and XAS data in the preceding paper.

Figure 2 plots the mean Ni-O distance for the compounds given in Fig. 1 as a function of the IOS.Ni. The mean Ni-O distance (R_{av}) for β -NiO(OH) and γ -"NiOOH" is calculated as $R_{av} = (N_1 R_1 + N_2 R_2) / (N_1 + N_2)$, where the R_1 and R_2 (Ni-O distances of the Ni-OH and Ni-O bonds) and the N_1 and N_2 (coordination numbers of the same) have been previously determined by EXAFS analyses [4, 5]. The resultant mean Ni-O distance decreases almost linearly as the IOS.Ni increases, except for that of γ -"NiOOH". An accurate mean Ni-O distance for γ -"NiOOH" is hard to obtain, because the measurement must be *in situ* and is affected by self-discharging. Moreover, two Ni-O bond distances are involved. The coordination number (N_1 and N_2) appear to be more sensitive to these experimental instabilities than the Ni-O distances (R_1 and R_2). A relatively large deviation from the linear relation for γ -"NiOOH" is likely due to this experimental difficulty. To remind as of this, we have added an arbitrary error bar for the γ -"NiOOH" datum. Unlike the Ni-O distances (Fig. 1), the linearity of the mean Ni-O distances (solid circles)

presented in Fig. 2 can be accounted for by either the OS.Ni=IOS.Ni or OS.Ni=2 premise, as shown below.

The OS.Ni=IOS.Ni premise for the $\text{Ni}(\text{OH})_2$ - $\text{NiO}(\text{OH})$ - NiO_2 series implies the $\text{Ni}^{2+}(\text{OH})_2^-$ - $\text{Ni}^{3+}\text{O}^{2-}(\text{OH})^-$ - $\text{Ni}^{4+}\text{O}^{2-}_2$ electronic structures. Accordingly, the Ni-OH and Ni-O bond lengths for them can be estimated as a sum of the ionic radii cited above. The resultant mean Ni-O distances gives the dashed lines in Fig. 2. On the other hand, the OS.Ni=2 premise for the $\text{NiO}_x(\text{OH})_{2-x}$ series implies, as discussed above, two constant Ni-O distances, R_{OH} for the Ni-OH bond and R_{O} for the Ni-O bond, independent of x . The mean Ni-O distance for the $\text{NiO}_x(\text{OH})_{2-x}$ is then given by

$$R_{\text{av}} = [xR_{\text{O}} + (2-x)R_{\text{OH}}]/2 = R_{\text{OH}} - (R_{\text{OH}} - R_{\text{O}})(\text{IOS.Ni} - 2)/2, \quad (1)$$

since $\text{IOS.Ni}=2+x$. This equation with the constants R_{OH} and R_{O} (obtained from Fig. 1) gives the solid line in Fig. 2. Thus, both the solid line (OS.Ni=2) and the dashed line (OS.Ni=IOS.Ni) reproduce nicely the experimental results (solid circles) in Fig. 2. However, as demonstrated above, the OS.Ni=IOS.Ni assignment for $\text{NiO}_x(\text{OH})_{2-x}$ does not reproduce the experimental results in Fig. 1.

III. Ni-O BOND LENGTH FOR THE NiO - LiNiO_2 - NiO_2 SERIES

The NiO, LiNiO_2 , and the $\text{IOS.Ni} \equiv 4$ compounds [i.e., $\text{Li}_{0.06}\text{NiO}_2$, $\text{KNiO}_2(\text{IO}_4)$, and BaNiO_3 , which are collectively referred to here as NiO_2] have six equal Ni-O bond lengths. Fig. 3 plots the Ni-O bond length for the NiO - LiNiO_2 - NiO_2 series as a function of the IOS.Ni, where the data for NiO [8] and LiNiO_2 [3, 9] are obtained from the cited references. The Ni-O bond length for this series again can be accounted for by either the OS.Ni=IOS.Ni or OS.Ni=2 premise, as will be seen at the end of this section.

The OS.Ni=IOS.Ni premise for this series implies the $\text{Ni}^{2+}\text{O}^{2-}$ - $\text{Li}^+\text{Ni}^{3+}\text{O}^{2-}_2$ - $\text{Ni}^{4+}\text{O}^{2-}_2$ electronic structures. Accordingly, the Ni-O bond length for this series can be estimated as a sum of the ionic radii of O^{2-} and Ni^{n+} , as we have done above. The resultant Ni-O bond lengths (dashed lines) are plotted in Fig. 3. Agreement between experiment and these estimates is excellent, supporting the conventional OS.Ni=IOS.Ni assignments and casting some doubt on the OS.Ni=2 assignments.

On the other hand, the OS.Ni=2 premise for this series implies the $\text{Ni}^{2+}\text{O}^{2-}$ - $\text{Li}^+\text{Ni}^{2+}[\text{O}^-\text{O}^{2-}]$ - $\text{Ni}^{2+}\text{O}^{2-}_2$ electronic structures. The $\text{Li}^+\text{Ni}^{2+}[\text{O}^-\text{O}^{2-}]$ structure would suggest two Ni-O bond lengths, one for $\text{Ni}^{2+}-\text{O}^-$ and another for $\text{Ni}^{2+}-\text{O}^{2-}$. However, experimental results [9] shows six equal Ni-O bond lengths for LiNiO_2 , indicating that the $\text{O}^-(2p^{-1})$ hole is delocalized. Therefore, the $\text{Li}^+\text{Ni}^{2+}\text{O}^{1.5-}_2$ structure is more realistic than the $\text{Li}^+\text{Ni}^{2+}[\text{O}^-\text{O}^{2-}]$ structure. We then have the series, $\text{Ni}^{2+}\text{O}^{2-}$ - $\text{Li}^+\text{Ni}^{2+}\text{O}^{1.5-}_2$ - $\text{Ni}^{2+}\text{O}^{2-}_2$. The observed Ni-O bond length ordering for this series is $\text{NiO} (2.09\text{\AA}) > \text{LiNiO}_2 (1.97\text{\AA}) > \text{NiO}_2 (1.85\text{\AA})$, that is,

$$\text{Ni}^{2+}-\text{O}^{2-} > \text{Ni}^{2+}-\text{O}^{1.5-} > \text{Ni}^{2+}-\text{O}^- \quad (2)$$

This ordering is obviously counter-intuitive, since a simple electrostatic argument suggests the reverse. The remainder of this section will be devoted to showing that this counter-intuitive ordering is caused by covalent shortening.

A. Ionic contribution

In this subsection, we shall estimate the ionic contributions to the $\text{Ni}^{2+}-\text{O}^{2-}$ bond in NiO and the $\text{Ni}^{2+}-\text{O}^-$ bond in NiO_2 . In both NiO and NiO_2 , the Ni atoms are coordinated octahedrally by oxygen atoms, but they have the NaCl and CdI_2 structures, respectively. According to Eq. (13-3) in Pauling's book [10], the Ni-O bond lengths R_1 and R_2 for NiO and NiO_2 can be approximated by

$$R_1 = (nB_1/A_1Q_1^2)^{1/(n-1)} \text{ and } R_2 = [nB_2/A_2(Q_2/2)^2]^{1/(n-1)}, \quad (3)$$

where the Born exponent n is around $n=8$ for the nickel oxides; A_1 ($=1.74756$) and A_2 ($=4.71$) are the Madelung constants for the NaCl and CdI₂ structures; B_1 and B_2 , the repulsive coefficients, should be similar ($B_1 \approx B_2$), since the number of coordinated anions to the Ni ion are the same; and Q_1 and Q_2 are the effective atomic charges on the Ni atoms, which are not necessarily equal to the OS.Ni. It is reasonable to assume $Q_1=Q_2$, because we are assuming the same OS.Ni ($=2$) for both NiO and NiO₂. Using the above relations, we obtain

$$R_2/R_1 = (4A_1/A_2)^{1/(n-1)} = 1.058. \quad (4)$$

Note that the R_2/R_1 ratio is independent of the B and Q coefficients. From the known Ni-O distance for NiO ($R_1=2.089\text{\AA}$) and the above R_2/R_1 ratio (4), we estimate the Ni-O distance for the *layered* NiO₂ ($R_2=2.210\text{\AA}$). In sum, the Ni²⁺-O⁻ bond length for NiO₂ is predicted to be longer than the Ni²⁺-O²⁻ bond length for NiO by 0.12\AA .

So far, we have implicitly assumed that the "size" of the O²⁻ and O⁻ ions are equal to each other by setting the repulsive coefficients equal, $B_1=B_2$. Here we estimate the relative sizes of the O²⁻ and O⁻ ions following Pauling's idea of univalent radii [10], which assumes that the ion size (R) is proportional to the mean radius (namely, $\langle\phi|r|\phi\rangle$) of the outermost orbital ϕ , (the $2p$ orbital for the O²⁻ and O⁻ ions). Using Slater's atomic wavefunctions and Burn's rules [11] for evaluating the screening constants S [i.e., $S(\text{O}^-)=4.40$; $S(\text{O}^{2-})=4.75$], we obtain

$$R(\text{O}^{2-})/R(\text{O}^-) = [Z-S(\text{O}^-)]/[Z-S(\text{O}^{2-})] = 1.108, \quad (5)$$

where Z ($=8$) is the nuclear charge of the oxygen atom. From the known O²⁻ radius $R(\text{O}^{2-})=1.26\text{\AA}$, we obtain $R(\text{O}^-)=1.14\text{\AA}$, which is shorter than the O²⁻ radius by 0.12\AA . Combining this and the result in the previous paragraph, we conclude that the Ni²⁺-O²⁻

bond length for NiO should be equal to the $\text{Ni}^{2+}\text{-O}^-$ bond length for the *layered* NiO_2 since the two effects, electrostatic and ion-size effects, are almost exactly canceled out by each other. To sum up, there is no ionic contribution to the Ni-O bond-length difference between the $\text{Ni}^{2+}\text{-O}^{2-}$ bond in NiO and the $\text{Ni}^{2+}\text{-O}^-$ bond in NiO_2 .

B. Covalent contribution

There is another mechanism that can alter the Ni-O bond length, i.e., covalent shortening. Since the O^- ion, unlike the O^{2-} ion, has an open-shell electronic structure, the $\text{Ni}^{2+}\text{-O}^-$ bond length in $\text{Ni}^{2+}\text{O}_2^-$ can be shorter than the $\text{Ni}^{2+}\text{-O}^{2-}$ bond length in $\text{Ni}^{2+}\text{O}^{2-}$, due to a covalent shortening. Below we estimate the magnitude of this covalent shortening.

In terms of the NiO_6 cluster approximation, $\text{Ni}^{2+}\text{O}_2^-$ and $\text{Ni}^{2+}\text{O}^{2-}$ corresponds to $\text{Ni}^{2+}\text{O}_6^-$ and $\text{Ni}^{2+}\text{O}_6^{2-}$, respectively. The electronic state, $\text{Ni}^{2+}\text{O}_6^-$, can be reached from the $\text{Ni}^{2+}\text{O}_6^{2-}$ by removing one electron from each oxygen atom, 6 electrons altogether. We therefore refer $\text{Ni}^{2+}\text{O}_6^{2-}$ and $\text{Ni}^{2+}\text{O}_6^-$ as the ground and ionized state, respectively. Figure 4 depicts the symmetric Ni-O stretching potential curves, $V^{(G)}(R)$ and $V^{(I)}(R)$, for the ground and ionized state, where R is the Ni-O distance, and $R_e^{(G)}$ and $R_e^{(I)}$ are the equilibrium Ni-O distances of the ground state and ionized state. The potential energy difference, $V^{(I)}(R) - V^{(G)}(R)$, is nothing but the ionization energy, $\text{IE}(R)$ [$\equiv V^{(I)}(R) - V^{(G)}(R)$]. As indicated in Fig. 4, the $\text{IE}(R)$ function varies slowly with R in comparison with the $V(R)$ functions. For a small variation of R , such as $R_e^{(G)} - R_e^{(I)} \approx 0.2 \text{ \AA}$, the derivative $d\text{IE}(R)/dR$, is almost invariant, whereas the $dV(R)/dR$ is almost a linear function of R . In the harmonic approximation, $V^{(G)}(R)$ is given by $k(R - R_e^{(G)})^2/2$, where k is a symmetric force constant.

The $R_e^{(I)}$ can be obtained from the following equations,

$$dV^{(1)}(R)/dR = dIE(R)/dR + dV^{(G)}(R)/dR = 0 \quad (6)$$

$$dV^{(G)}(R)/dR = k(R - R_e^{(G)}) \quad (7)$$

$$dIE(R)/dR = -6 d\epsilon_{t_{1u}}(R)/dR \quad (8)$$

where $\epsilon_{t_{1u}}$ is the orbital energy of the highest orbital arising from the ligand oxygen 2p orbitals, which has the t_{1u} symmetry of the octahedral NiO_6 system. This t_{1u} orbital is fully occupied in the ground state containing 6 electrons, but is emptied in the ionized state. Note that the $IE(R)$ is approximated by $-6\epsilon_{t_{1u}}(R)$, as usual. Equations (6)-(8) give

$$R_e^{(1)} - R_e^{(G)} = (6/k) d\epsilon_{t_{1u}}(R)/dR, \quad (9)$$

where $d\epsilon_{t_{1u}}(R)/dR$ is almost invariant and hence will be evaluated at $R=R_e^{(G)}$. Since the t_{1u} orbital is antibonding with respect to the Ni-O bond, the derivative $d\epsilon_{t_{1u}}/dR$ will be negative and thereby Eq. (8) will predict $R_e^{(1)} < R_e^{(G)}$ or $Ni^{2+}-O^- < Ni^{2+}-O^{2-}$ bond-length ordering, i.e., a covalent shortening.

We now estimate k and $d\epsilon_{t_{1u}}(R)/dR$ and evaluate $R_e^{(1)} - R_e^{(G)}$. The symmetric stretching force constant k ($=2.321 \times 10^5$ [dyne/cm]) is estimated from the Ni-O symmetric stretching frequency ($\omega_e = 560$ cm^{-1}), which has been observed for the charged NOE [12], using the relation $k = 5.8883 \times 10^{-2} \mu \omega_e^2$ [dyne/cm], where μ is the reduced mass of Ni and O in amu. The $d\epsilon_{t_{1u}}(R)/dR$ at $R=R_e^{(G)}$ can be calculated rather accurately by any method which includes the overlap integral between the metal and the ligand oxygen atomic orbitals. We have employed the extended Hückel method with no charge iteration and obtained $d\epsilon_{t_{1u}}(R)/dR = -0.8401 \times 10^{-4}$ [erg/cm] at the $R=R_e^{(G)} = 2.09 \text{ \AA}$. Substituting this $d\epsilon_{t_{1u}}(R)/dR$ and the k value into Eq. (4), we have $R_e^{(1)} - R_e^{(G)} = -0.22 \text{ \AA}$. This is the covalent contribution to the Ni-O bond length. Since there is no ionic contribution, we predict that the $Ni^{2+}-O^-$ bond length in NiO_2 is shorter than the $Ni^{2+}-O^{2-}$ bond length in NiO by 0.22 \AA due to covalent shortening. This prediction is shown in Fig. 3 by the solid line.

Figure 3 shows that both the solid line (i.e., covalent shortening assuming the OS.Ni=2) and the dashed line (i.e., the sum of ionic radii assuming the OS.Ni=IOS.Ni) reproduce excellently the experimental results. However, the OS.Ni=2 assignment is more consistent with the other data. Therefore, we can attribute the decrease of the Ni-O bond length in the NiO-LiNiO₂-NiO₂ series (Fig. 3) to a covalent shortening.

IV. JAHN-TELLER DISTORTION

the absence or presence of a Jahn-Teller distortion (JTD) for the IOS.Ni=3 compounds [i.e., LiNiO₂ and β -NiO(OH)] provides further evidence regarding our question. The ground states of the Ni²⁺ and Ni³⁺ ions in the NiO₆ octahedron are expected to be Ni²⁺(t_{2g}⁶e_g²)³A_{2g} and Ni³⁺(t_{2g}⁶e_g¹)²E_g. A strong JTD is expected for the degenerate E state of the Ni³⁺ ion, but not for the non-degenerate A state of the Ni²⁺ ion. For LiNiO₂, no JTD of the NiO₆ octahedron has been observed [3, 9, 13], suggesting the OS.Ni=2 assignment.

β -NiO(OH) also appears to lack a JTD. The α -AlO(OH) and α -MnO(OH) compounds are known to have a layered structure similar to that of β -NiO(OH) [14]. α -MnO(OH), having high-spin Mn³⁺(t₂³e¹), shows a strong JTD, whereas α -AlO(OH), having a closed-shell ion, Al³⁺, lacks a JTD. As shown in Fig. 5, the Mn-OH and Mn-O bond lengths are clearly split due to the JTD, whereas the Al-OH and Al-O bond lengths are barely split. The M-OH and M-O data for M=Al and Mn are obtained from Wells' book [14]. A splitting of the Ni-OH and Ni-O bond lengths has not been observed for β -NiO(OH) [4]. This absence of a JTD for β -NiO(OH) favors the Ni²⁺ assignment over the conventional Ni³⁺ assignment.

V. SUMMARY OF EVIDENCE

Table I summarizes, for each compound, the supporting evidence cited for the OS.Ni=2 vs. OS.Ni=IOS.Ni assignments. This evidence is described with shorthand notations, which are explained in the footnotes of the table. The XPS and XAS evidence are presented in the preceding paper and the geometric evidence is given in this paper. The remainder of the evidence listed in the table, (i.e., "Magnetic moment" and "Oxygen K-edge XAS") can be found in the references given there.

Let us first evaluate the evidence listed for the LiNiO₂ compound. The oxygen K-edge XAS spectrum provides direct evidence for the presence of the O⁻ anion [15]. This spectrum and the absence of a Jahn-Teller distortion would be very hard to account for with the Li⁺Ni³⁺O²⁻₂ electronic structure, if not impossible. On the other hand, the observed magnetic moment, which might strongly suggest Ni³⁺ [13, 16], can be accounted for with a strong antiferromagnetic coupling between the Ni²⁺ and O⁻ spins [15] based on the OS.Ni=2 assumption. Furthermore, the observed Ni-O bond length, which agrees amazingly well with a sum of the ionic radii of Ni³⁺ and O²⁻ ions, can also be accounted for by a covalent shortening as demonstrated in this work. In addition, the XPS main-peak energy and satellite energy separation strongly suggest the OS.Ni=2 assignment. For LiNiO₂, therefore, the evidence for OS.Ni=2 outweighs that for OS.Ni=3.

The LiNiO₂ case shows that the magnetic-moment and Ni-O (ionic radii) evidence does not prove the OS.Ni=NOS.Ni assignment for LiNiO₂. Therefore, this evidence likely does not prove the OS.Ni=NOS.Ni for the other compounds either. But these are the only two points supporting the OS.Ni=NOS.Ni assignments as seen in the table. On the other hand, there is a preponderance of evidence for the OS.Ni=2 assignment, especially for β-NiO(OH) and γ-NiOOH. Further, the Ni-O (JTD) evidence for β-NiO(OH) and the XPS (line width) evidence for γ-NiOOH would be very hard to account for assuming the

OS.Ni=IOS.Ni premise. More importantly, the OS.Ni=2 assignment can *consistently* account for *all* of the experimental data available. In sum, we propose the OS.Ni=2 assignment for all compounds listed in Table I.

Although we have concluded that the OS.Ni=2 for all the compounds listed in Table I, we certainly do not suggest OS.Ni=2 for every Ni oxide. Our choice of materials is rather restricted, since the Ni site in our materials is octahedrally coordinated by oxygen atoms and all of our materials, except $\text{KNiO}_2(\text{IO}_4)$ and BaNiO_3 , have a layered structure. (BaNiO_3 has a one-dimensional chain structure [6]; the structure of $\text{KNiO}_2(\text{IO}_4)$ is unknown but likely has a layered structure.) Under these restrictions, we have found that the OS.Ni is 2.

VI. IMPLICATIONS

Various physical properties, such as the optical, magnetic, and electrical properties, of the nickel oxide electrode (NOE) and related nickel compounds have been discussed on the premise of the OS.Ni=IOS.Ni assignment, which is currently accepted by a majority of workers in the field. To our knowledge, only exception to this is that Sawatzky's group has attempted to account for the magnetic property of Li-substituted NiO [i.e., $\text{Li}_y\text{Ni}_{2-y}$ for $y=0-1$] based on the premise of OS.Ni=2 [15]. Now that the OS.Ni=2 assignment is established here, previous explanations based on the OS.Ni=IOS.Ni assignment must be re-examined in light of this new finding. Below we simply list, without proof, the implications of the OS.Ni=2 assignment on the physical properties.

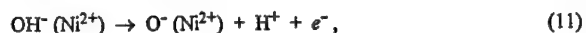
A. Charging reactions

The overall reaction in the charging of the nickel oxide electrode (NOE) is, in the

simplest form,

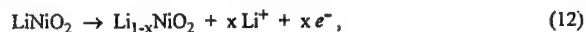


where the H^+ ion and electron, e^- , move outside of the NOE material. Relying on the ionic picture of the OS.Ni=IOS.Ni , this reaction has been viewed as oxidation of Ni^{2+} into the higher oxidation states, Ni^{3+} and/or Ni^{4+} . According to the new assignment of OS.Ni=2 , it should be viewed as oxidation of O^{2-} , namely,

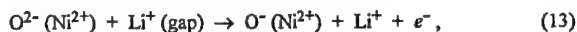


where (Ni^{2+}) reminds us that the OH^- and O^- anions are attached to the Ni^{2+} ion, which remains unchanged.

Similarly, the charging reaction of another battery material LiNiO_2 is given by



where the Li^+ ion and electron, e^- , leave the LiNiO_2 material. Again relying on OS.Ni=IOS.Ni , this reaction has been viewed as oxidation of the Ni^{3+} into Ni^{4+} [3]. According to the OS.Ni=2 assignment, it should be viewed as oxidation of the O^{2-} ,



where the $\text{Li}^+ (\text{gap})$ means that the Li^+ ion is in the van der Waals gap of the NiO_2 slabs.

B. Color change

Upon charging NOE $[\text{Ni(OH)}_2]$, its color (pale green) turns to black, the apparent color of $\beta\text{-NiO(OH)}$ and $\gamma\text{-NiOOH}$ [17, 18]. Assuming OS.Ni=IOS.Ni , this coloration has been attributed to the $2p(\text{O}^{2-}) \rightarrow 3d(\text{Ni}^{3+})$, oxygen-to-metal, charge transfer in the charged NOE [18]. On the other hand, the OS.Ni=2 assignment suggests another

possibility of $3d(Ni^{2+}) \rightarrow 2p(O^-)$, metal-to-oxygen, charge transfer for the strong optical absorption in the charged NOE. We favor the latter, the metal-to-oxygen charge transfer, based on the argument given in the APPENDIX A.

C. Electron conduction

On the premise of $OS.Ni=IOS.Ni$, electron conduction in the NOE has to take place in the Ni layer for both the charged and discharged NOE, because every oxygen ion has a closed electronic structure of $O^{2-}(2p^6)$. On the other hand, the $OS.Ni=2$ assignment suggests electronic conduction in the oxygen layer for the charged NOE though the holes of the $O^-(2p^5)$ ions. Electron conduction in the oxygen layer should be more efficient than that in the nickel layer, because the shortest O-O and Ni-Ni distances are equal in the NOE material but the size of the oxygen 2p orbital is much larger than that of the nickel 3d orbital. When enough O^- ions are generated, the oxygen layer could become a metallic conductor.

It is known that upon charging the NOE, the electrical conductivity increase three orders of magnitudes [17]. Also, upon overcharging, sometimes, a layer with metallic luster has been observed [19, 20]. These may be due to the electronic conduction in the oxygen layer.

D. Crystal structure

It is known that the $Li_{0.5}NiO_2$ compound has at least two polymorphs, the *layered* $Li_{0.5}NiO_2$ and cubic *spinel* $Li[Ni_2]O_4$ structures [13]. In both structures, Ni is octahedrally coordinated by oxygen atoms, but the spinel $Li[Ni_2]O_4$ has a 3D-networked structure whereas the layered $Li_{0.5}NiO_2$ has a 2D-layered structure. The layered $Li_{0.5}NiO_2$ is metastable and can be converted into the more stable $Li[Ni_2]O_4$ by heating around 250°C [13]. In this conversion process, one quarter of the Ni from the NiO_2 slabs

of the layered $\text{Li}_{0.5}\text{NiO}_2$ move into the van der Waals gaps so as to maximize the electrostatic Madelung energy, resulting in a 3D-networked structure of NiO_2 . By analogy, we speculate that the layered $\text{NiO}_x(\text{OH})_{2-x}$ is also metastable and can be converted by heating at a high temperature into a more stable 3D-networked structure.

We believe that NiO_2 also has two or more polymorphs, the layered NiO_2 and the commercial NiO_2 material which was supplied by Aldrich Chemical Co. (Milwaukee, WI). The layered NiO_2 may be metastable and can be converted into the more stable commercial NiO_2 material by heating. We suspect that the commercial NiO_2 material has an amorphous 3D-networked structure, in which the Ni atoms are coordinated by oxygen anions in a tetrahedral geometry as well as octahedral. Evidence for the polymorphism in the NiO_2 compound is given in APPENDIX B.

We speculate that a close relationship exists between the crystal structure (2D-layered and 3D-networked) and the electronic structure (covalent and ionic). The more stable 3D-networked polymorph has more ionic bonding, or $\text{OS.Ni}=\text{IOS.Ni}$, whereas the layered polymorph has more covalent bonding, or $\text{OS.Ni}=2$.

ACKNOWLEDGMENTS

We gratefully acknowledge the support of the Office of Naval research.

APPENDIX A: Optical absorption of NOE

The absorption spectra on thin ($< 200\text{\AA}$) film $\text{Ni}(\text{OH})_2$ electrodes have been measured with *in situ* reflection [21] and transmission [18] spectroscopies. The results from the two techniques agree with each other and reveal the following characteristics:

(i) There is virtually no absorption for the fully reduced NOE (i.e., $\text{Ni}(\text{OH})_2$) in the 300-900nm region. In contrast, the spectra of the fully charged NOE (i.e., $\gamma\text{-NiOOH}$) in this region shows a strong absorption which cannot be attributable to "forbidden" transitions such as *d-d* transitions.

(ii) The 2-5eV absorption band consists of two features, a broad and strong band peaked around 3.0eV and a sharp but weak band peaked around 4.1eV.

(iii) The spectra at intermediate states of charge show that the absorbency of the 2-5eV band is proportional to the charging with no new absorption bands appearing.

Below we examine the OS.Ni=2 and OS.Ni=3.5 alternate assignments for $\gamma\text{-NiOOH}$ based on these observed characteristics (i)-(iii).

Based on the OS.Ni=2 premise, we can attribute the observed optical bands to metal-to-oxygen charge transfer, $3d(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$, more precisely, $e_g(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ and $t_{2g}(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ transitions from the ground state $[\text{Ni}^{2+}(t_{2g}^6 e_g^2) \text{O}^{2-}(2p^5) \dots]$ of a charged NOE. The crystal field splitting of the *d* orbitals in an octahedral field Δ_o (i.e., t_{2g} and e_g orbital energy difference) for the first transition metals are known to be in the range of 0.9-1.6eV for divalent metals and 1.7-3.1eV for trivalent metals [22]. Δ_o for NiO solid is usually assumed to be around 1eV [23]. Band peaks due to the $e_g(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ and $t_{2g}(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ transitions, therefore, will be separated by about 1eV. Furthermore, we expect that the $e_g \rightarrow 2p$ band is broader and stronger than the $t_{2g} \rightarrow 2p$ band. In short, the $e_g \rightarrow 2p$ and $t_{2g} \rightarrow 2p$ bands will have the same characteristics as the 3.0 eV and 4.1 eV

bands described in (ii).

The characteristics of (i) and (iii) can also be accounted for quite naturally by the $3d(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ assignment. The charge-transfer transitions appear only after 2p holes are created in the O^{2-} ions producing O^- ions. This immediately explains (i) why the optical absorption band appears only after the NOE is charged. The charging of the NOE in the OS.Ni=2 assignment increases the number of O^- ions but not the OS.Ni value. Consequently the absorbency, i.e., the amplitude of the $3d(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ bands, increases proportionally to the charging, but the shape of the absorption bands is independent of the charging, since the same $3d(\text{Ni}^{2+}) \rightarrow 2p(\text{O}^-)$ transitions are involved throughout the charging as observed (iii). To sum up, the observations (i)-(iii) are consistent with the OS.Ni=2 assignment for γ -NiOOH.

On the other hand, the explanations of (i)-(iii) based on the OS.Ni=3.5 premise for γ -NiOOH is problematic. The 2-5 eV band has been previously [18, 21] attributed to the oxygen-to-metal charge transfer, $2p(\text{O}^{2-}) \rightarrow 3d(\text{Ni}^{3+})$, on the analogy of the absorption spectrum of Ni^{3+} ions in corundum (Al_2O_3), which exhibits two relatively narrow (FWHM ≈ 0.7 eV) and almost equal amplitude bands around 3.0 and 4.5 eV [24]. The OS.Ni=3.5 premise, however, implies the presence of Ni^{4+} ions along with Ni^{3+} ions. The $2p(\text{O}^{2-}) \rightarrow 3d(\text{Ni}^{4+})$ transition should produce another band at lower energy but such an additional band has not been observed down to 1.5 eV. This absence of the $2p(\text{O}^{2-}) \rightarrow 3d(\text{Ni}^{4+})$ absorption band has to be explained. Another difficulty is (iii), which implies the presence of only one colored (homogeneous) species with OS.Ni=3.5. Recent studies [3, 25] on the similar system Li_xNiO_2 ($x=0-1$), however, indicates that Li_xNiO_2 goes through at least three phases on the oxidation-reduction process. In sum, the OS.Ni=3.5 assignment for γ -NiOOH may not be inconsistent with (i)-(iii), but unusual explanations are required to account for them.

APPENDIX B: Polymorphism of NiO_2 compound

Mansour's group has studied commercial NiO_2 material (supplied by Aldrich Chemical Co., Milwaukee, WI) by X-ray diffraction (XRD) and X-ray absorption [5, 26]. The XRD spectrum shows broad bands indicative of highly disordered or amorphous materials. [Similar broad XRD bands have been reported by Melendres's group [27] for NiO_2 supplied by Alfa-Ventron (Danvers, MA) and also for NiO_2 samples prepared chemically following either the Bro and Cogley [28] or Nakagawa [29] procedure.] In contrast, the XRD spectrum for layered $\text{Li}_{0.06}\text{NiO}_2$ shows sharp lines indicative of crystalline materials [3].

EXAFS analyses for the commercial NiO_2 [5] yielded two Ni-O bond lengths (1.91 and 2.06 Å) with a weighted average of 1.99 Å, which is much longer than that (1.852 Å) of the layered NiO_2 . Moreover, the latter has only one Ni-O bond length. The pre-edge peak of the layered NiO_2 is expected to be sharp and shifted about 2 eV (judging from its Ni-O bond length) to higher energy relative to that of NiO . The pre-edge peak observed for the commercial NiO_2 material is, however, broad and shifted only by ≈ 0.7 eV. The K-edge shift of the commercial NiO_2 material is also small (≈ 1.3 eV) in comparison with that (3.0 eV) of the layered NiO_2 compound.

All these differences between the layered NiO_2 and the commercial NiO_2 material imply that they are indeed different polymorphs.

REFERENCES

- [1] J. McBreen, in *Modern Aspects of Electrochemistry*, eds. R.E. White, J. O'M. Bockris, and B.E. Conway (Plenum Press, New York, 1990), Vol. 21, p. 29.
- [2] C. Delmas, *Mat. Res. Soc. Symp. Proc.* **293**, 15 (1993).
- [3] T. Ohzuku, A. Ueda, and M. Nagayama, *J. Electrochem. Soc.* **140**, 1862 (1993).
- [4] K.I. Pandya, W.E. O'Grady, D.A. Corrigan, J. McBreen, and R.W. Hoffman, *J. Phys. Chem.* **94**, 21 (1990); K.I. Pandya, P.W. Hoffman, J. McBreen, and W.E. O'Grady, *J. Electrochem. Soc.* **137**, 383 (1990).
- [5] A.N. Mansour and C.A. Melendres, *Physica B* **208&209**, 583 (1995).
- [6] Y. Takeda, F. Kanamaru, M. Shimada, and M. Koizumi, *Acta Cryst.* **B 32**, 2464 (1976).
- [7] R.D. Shannon, *Acta Cryst. A* **32**, 751 (1976).
- [8] R.W.G. Wyckoff, *Crystal Structures* (Interscience Publishers, New York, 1964) Second Edition, Vol. 1, Table III-1, p. 89.
- [9] R.W.G. Wyckoff, *Crystal Structures* (Interscience Publishers, New York, 1964) Second Edition, Vol. 2, Table VI-3.
- [10] L. Pauling, *The Nature of the Chemical Bond* (Cornell University Press, Ithaca, 1960).
- [11] G. Burns, *J. Chem. Phys.* **41**, 1521 (1964).
- [12] J. Desilvestro, D.A. Corrigan, and M.J. Weaver, *J. Electrochem. Soc.* **135**, 885 (1988).
- [13] G. Dutta, A. Manthiram, J.B. Goodenough, J.-C. Grenier, *J. Solid State Chem.* **96**, 123 (1992).
- [14] A.F. Wells, *Structural Inorganic Chemistry* (Clarendon Press, Oxford, 1984), 5th edition, Table 14.2, p. 641.

- [15] P. Kuiper, G. Kruizinga, J. Ghijsen, and G.A. Sawatzky, *Phys. Rev. Lett.* **62**, 221 (1989); J. van Elp, H. Eskes, P. Kuiper, and G.A. Sawatzky, *Phys. Rev. B* **45**, 1612 (1992).
- [16] J.P. Kemp, P.A. Cox, and J.W. Hodby, *J. Phys. Condens. Matter* **2**, 6699 (1990).
- [17] M.J. Natan, D. Bélanger, M.K. Carpenter, and M.S. Wrighton, *J. Phys. Chem.* **91**, 1834 (1987).
- [18] D.A. Corrigan and S.L. Knight, *J. Electrochem. Soc.* **136**, 613 (1989).
- [19] G.W.D. Briggs and M. Fleischmann, *Trans. Faraday Soc.* **67**, 2397 (1971).
- [20] D. Tuomi, *J. Electrochem. Soc.* **112**, 1 (1965).
- [21] J.D.E. McIntyre and D.M. Kolb, *Symp. Faraday Soc.* **4**, 99 (1970).
- [22] F.A. Cotton and G. Wilkinson, *Advanced Inorganic Chemistry* (Interscience Publishers, New York, 1972), Third Edition, p. 577.
- [23] A. Fujimori and F. Minami, *Phys. Rev. B* **30**, 957 (1984).
- [24] H.H. Tippins, *Phys. Rev. B* **1**, 126 (1970).
- [25] W. Li, J.N. Reimers, and J.R. Dahn, *Solid State Ionics* **67**, 123 (1993).
- [26] A.N. Mansour, C.A. Melendres, M. Pankuch, and R.A. Brizzolara, *J. Electrochem. Soc.* **141**, L69 (1994).
- [27] C.A. Melendres, W. Paden, B. Tani, and W. Walczak, *J. Electrochem. Soc.* **134**, 762 (1987).
- [28] P. Bro and D. Cogley, *J. Electrochem. Soc.* **113**, 521 (1966).
- [29] K. Nakagawa, R. Konaka, and T. Nakata, *J. Amer. Chem. Soc.* **27**, 1597 (1962).

Table I. Summary of supporting evidence for the OS.Ni=2 and OS.Ni=IOS.Ni assignments for each compound, where OS.Ni stands for "oxidation state of nickel" and the value for the IOS.Ni (ionic OS.Ni) is given in the table.

Compound	IOS.Ni	Supporting evidence for	
		OS.Ni = 2	OS.Ni = IOS.Ni
LiNiO ₂	3	XPS (main peak) XPS (satellite) Ni-O (JTD) Ni-O (covalent) Oxygen K-edge XAS [15]	Magnetic moment [13, 16] Ni-O (ionic radii)
β-NiO(OH)	3	XAS (K-edge) Ni-O (JTD) Ni-O (constant)	Magnetic moment [2] Ni-O (ionic radii)
γ-"NiOOH"	3.5	XPS (main) XPS (satellite) XPS (line width) XAS (white line) XAS (k-edge) Ni-O (constant)	Magnetic moment [2] Ni-O (ionic radii)
KNiO ₂ (IO ₄)	4	XPS (main) XPS (satellite) XAS (pre-edge peak) Ni-O (covalent)	Ni-O (ionic radii)
BaNiO ₃	4	XAS (K-edge) Ni-O (covalent)	Ni-O (ionic radii)

Ni-O (JTD) = absence of Jahn-Teller distortion.

Ni-O (constant) = constant Ni-OH and Ni-O bond lengths.

Ni-O (ionic radii) = sum of ionic radii.

Ni-O (covalent) = covalent shortening.

XPS = x-ray photoemission spectroscopy

XAS = x-ray absorption spectroscopy.

FIGURE CAPTIONS

FIG. 1. Ni-O distances of the Ni-OH (open circle) and Ni-O (closed circle) bonds observed for $\text{NiO}_x(\text{OH})_{2-x}$ [$x=0, 1$, and 1.5] as a function of the IOS.Ni along with the Ni-O bond lengths for the IOS.Ni=4 compounds, $\text{Li}_{0.06}\text{NiO}_2$, $\text{KNiO}_2(\text{IO}_4)$, and BaNiO_3 . The dashed and dotted lines are respectively the variations of the Ni-OH and Ni-O bond lengths predicted on the OS.Ni=IOS.Ni premise. The solid lines are those predicted on the OS.Ni=2 premise.

FIG. 2. Mean Ni-O distance for the compounds in Fig. 1 plotted as a function of the IOS.Ni. A large deviation from the linear relationship for γ -"NiOOH" is due to experimental difficulty (see the text). The dashed line is an averaged sum of the ionic radii assuming the OS.Ni=IOS.Ni. The solid line is the variation predicted assuming the OS.Ni=2.

FIG. 3. Ni-O bond length for the Ni-LiNiO₂-NiO₂ series plotted as a function of the IOS.Ni. The dashed line is obtained as a sum of the ionic radii assuming OS.Ni=IOS.Ni. The solid line is the variation of the Ni-O bond length predicted assuming OS.Ni=2. It is shown in the text that this variation is due to a covalent shortening.

FIG. 4. The relationship between the symmetric stretching potential curves, $V^{(0)}(R)$ and $V^{(1)}(R)$, for the ground $\text{Ni}^{2+}\text{O}_6^{2-}$ state and the ionized $\text{Ni}^{2+}\text{O}_6^-$ state. The difference, $V^{(1)}(R)-V^{(0)}(R)$, gives the ionization energy, $\text{IE}(R)$, which varies slowly with R in comparison with the $V^{(0)}(R)$ and $V^{(1)}(R)$ functions.

FIG. 5. The M-O distances of the M-OH (thin solid line) and M-O (thick solid line) bonds in $\text{MO}(\text{OH})$ for $\text{M}=\text{Al}$, Ni , and Mn . The M-OH and M-O bond lengths for $\text{MnO}(\text{OH})$ are splitted due to the Jahn-Teller distortion.

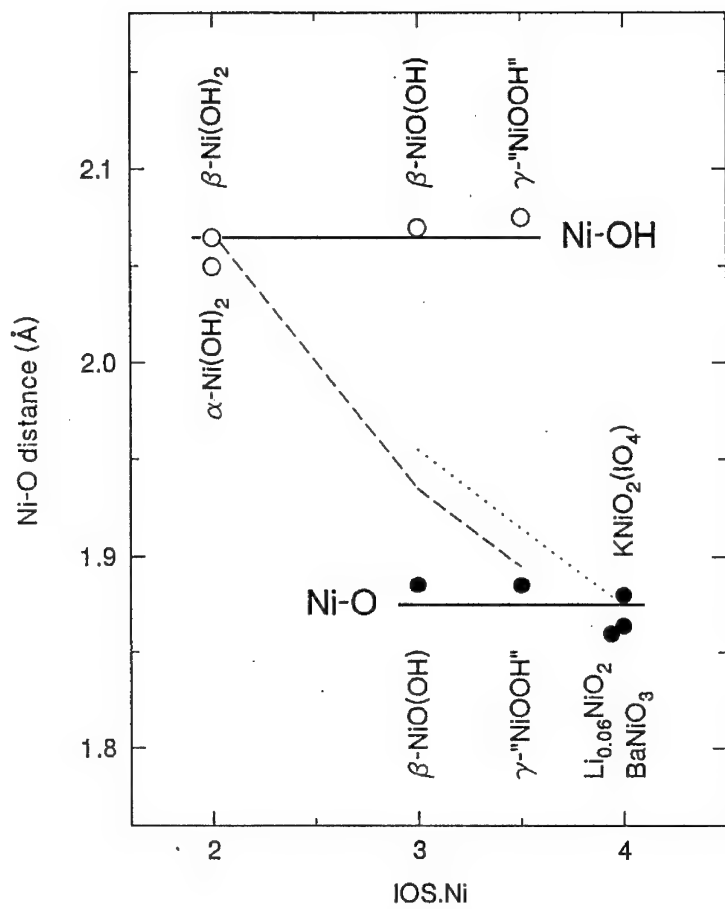


Fig. 1

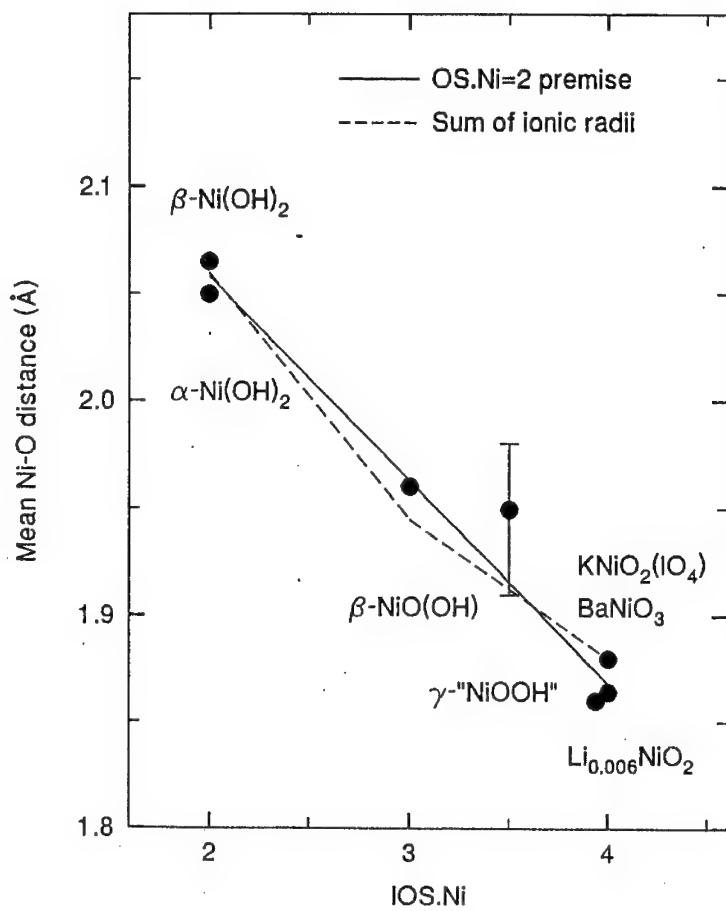


Fig. 2

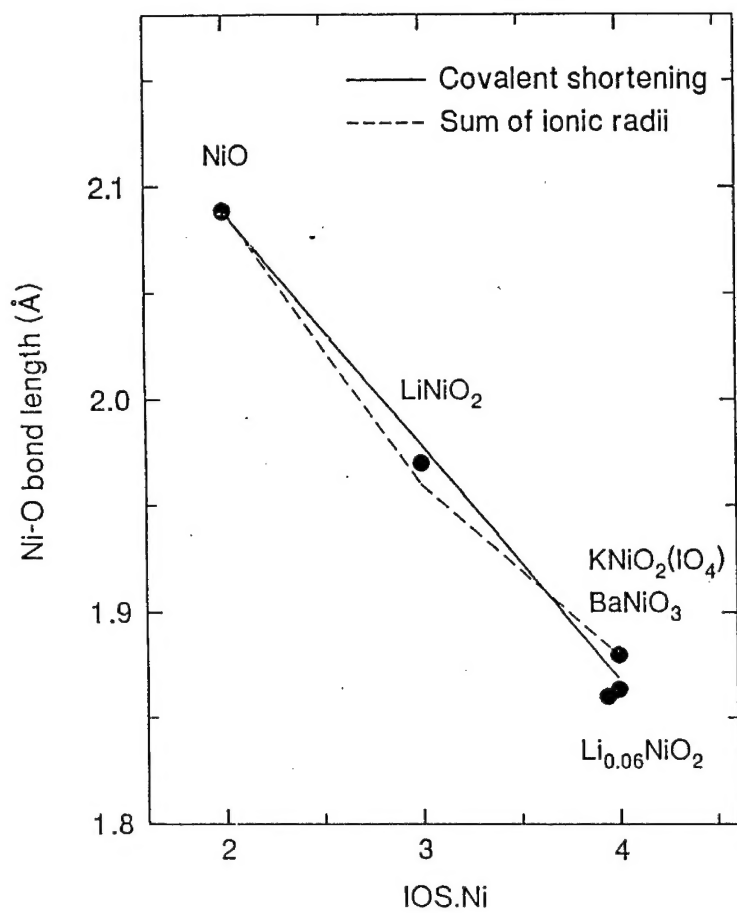


Fig. 3

Symmetric stretching potential energy $V(R)$

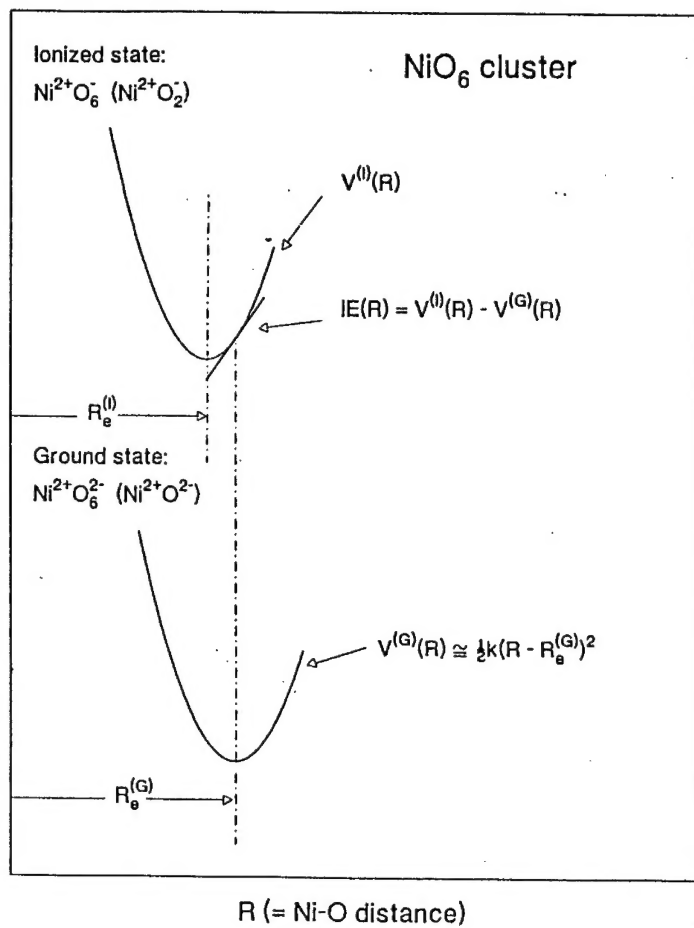


Fig. 4

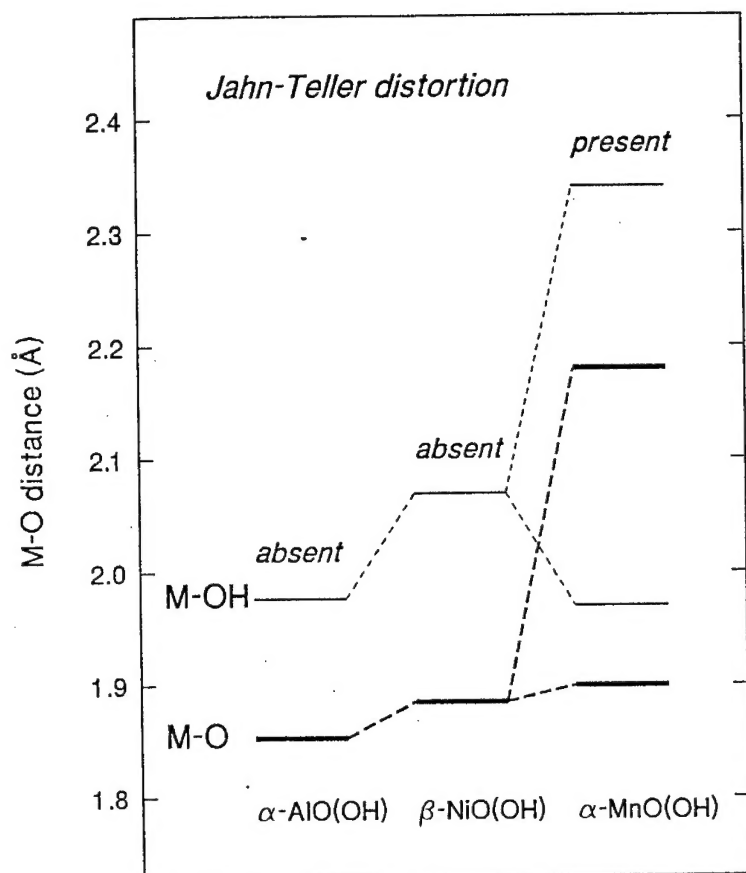


Fig. 5

Technical Report Distribution List

Dr. Robert J. Nowak (1)*
ONR 331
800 N. Quincy St.
Arlington, VA 22217-5660

Defense Technical Information Ctr (2) **
Building 5, Cameron Station
Alexandria, VA 22314

Dr. James S. Murday (1)
Chemistry Division, NRL 6100
Naval Research Laboratory
Washington, DC 20375-5660

Dr. John Fischer (1)
Chemistry Division, Code 385
NAWCWD - China Lake
China Lake, CA 93555-6001

Dr. Peter Seligman (1)
NCCOSC - NRRAD
San Diego, CA 92152-5000

Dr. James A. Gucinski (1)
NSWC Code 609
300 Highway 361
Crane, IN 47522-5001

Mr. Christopher Egan (1)
Naval Undersea Warfare Center
Division Newport
1176 Howell St.
Newport, RI 02841-1708

Dr. Carl Mueller
Naval Surface Warfare Center - White Oak
Code R36
10901 New Hampshire Ave.
Silver Spring, MD 20903-5640

* Number of copies required